Introduction

Reaction between cations and anions will be examined in this lab. The resulting precipitate is indicative of the reaction that occured.

In our lab a lead nitrate solution, $Pb(NO_3)_2$ will be mixed with a solution of potassium chromate, K_2CrO_4 , a yellow precipitate will forms. This precipitate must be a combination of K^+ or Pb^{2+} cation and CrO_4^{2-} or NO_3^{2-} anion. Since most potassium salts and nitrates are soluble in water, $PbCrO_4$ must be the precipitate. The equation for this reaction is:

Pb²⁺ + CrO₄ ²⁻ -> PbCrO₄